

Printed pages: 2	BOP244
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**BPHARM
(SEM II) THEORY EXAMINATION 2014-15
PHARMACEUTICAL ANALYSIS-II**

Time allowed: 3 Hrs.

Max Marks: 70

Section- A

1. Attempt all questions, each carries 2 marks. 2x10 = 20

- a. Write properties of solvents in used in non aqueous titrations .
- b. Write about differentiating solvent
- c. Define indicator electrode.
- d. What do you mean Complexometric titration
- e. Write down any four functions of salt bridge
- f. Define chromatography.
- g. What are the limitations of Kjeldahl method?
- h. Write down the advantages of glass electrode.
- i. Define Non-aqueous titrations
- j. Write down any four indicator used in non-aqueous titration.

Section- B

2. Attempt any four parts. Each carries equal marks. 5x4 = 20

- a. Explain non- aqueous solvents with suitable examples.
- b. Discuss applications of Complexometric titrations.
- c. How will you standardise 0.1N perchloric acid
- d. Write short notes on reference electrode.

- e. Discuss the determination of pH using glass electrode.

Section- C

3. Answer any four questions. 7.5x4 = 30

- a. Write down application of polarography.
- b. Write down principle of conductometric titration. Write down advantages of non-aqueous titration over aqueous titration. Discuss in brief application of non- aqueous titrations.
- c. Explain what is complexometric titration? Explain the role of masking and demasking agents in complexometric titrations?
- d. Discuss the principle involved in HPLC chromatography. How Write application of HPLC over TLC.
- e. Discuss the general principle instrumentation & application of Amperometry.
- f. Write down principle , instrumentation and application of potentiometry.